

SUPPORT FOR THE AMENDMENTS

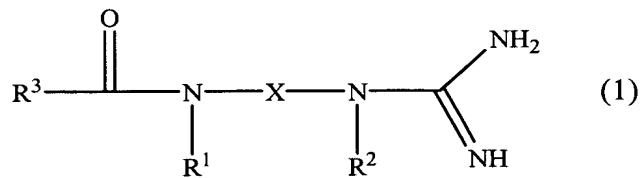
Applicants have amended Claim 1 for clarity and to incorporate the limitation of Claim 2. Accordingly, support for amended Claim 1 can be found in Claims 1 and 2, as originally filed. Claims 6 and 7 have been amended for clarity. Support for amended Claims 6 and 7 can be found in the same claims, as originally filed. Applicants have also added new Claims 12 to 31. Support for new Claims 12, 14, 15, 17, 18, 20, 22, 24, 25, 27, 28, 30, and 31 can be found in Claim 2, as originally filed. Support for new Claim 13 can be found on page 10, lines 8-13, of the specification. Support for new Claim 16 can be found on page 10, line 19, to page 11, line 8, of the specification. Support for new Claim 19 can be found on page 11, lines 18-23, of the specification.

No new matter has been added. Claims 1, 2, 6, 7, and 12-31 are active in this application.

REMARKS/ARGUMENTS

Present Claims 1, 2, 6, 7, and 12-31 relate to liquid compositions which comprise:

(A) at least one amide group-containing guanidine compound represented by formula (1):



wherein: R¹ and R² are independently a hydrogen atom, a linear or branched alkyl group having one to 4 carbon atoms, or a linear or branched alkenyl group having one to 4 carbon atoms and may be the same or different; R³ represents a linear or branched alkyl group

having one to 22 carbon atoms or a linear or branched alkenyl group having one to 22 carbon atoms; and X represents a linear or branched alkylene group having one to 10 carbon atoms or a linear or branched alkenylene group having one to 10 carbon atoms or a salt thereof; and

(B) at least one inorganic salt, which is selected from the group consisting of sodium chloride, potassium chloride, and mixtures thereof,

wherein said at least one amide-group containing guanidine compound (A) and said at least one inorganic salt (B) are present in a weight ratio of said at least one amide-group containing guanidine compound (A) to said at least one inorganic salt (B) of 1:0.5 to 1:20.

The inventors have discovered that the present compositions exhibit an enhanced stability. The cited references do not disclose or suggest the presently claimed compositions or the benefits provided thereby. Accordingly, these references cannot affect the patentability of the present claims.

The rejection of Claims 1, 2, 6, 7, and 11 under 35 U.S.C. § 102(b) in view of JP2001-18775 (Lion '775) the Merck Index, 10th Ed., pp. 1100, 1232, and 1233 (1983) (Merck) is respectfully traversed. At the outset, it is noted that the Examiner appears to recognize that Lion '775 does not disclose any compositions which contain both (A) a guanidine compound as recited in the present claims and (B) either sodium or potassium chloride in the claimed ratio. Rather, on page 3 of the Office Action, it is asserted that Lion '775 discloses the addition of "an electrolyte" for stabilization. Merck is cited for the proposition that sodium and potassium chloride are electrolytes.

However, even if the characterization of the cited references is correct, their combined disclosures neither anticipate nor make obvious the present claims.

In particular, it is well known in the art that the term “electrolyte” is a substance that dissociates into ions when dissolved in water.¹ Thus, the term “electrolyte” encompasses a great number of inorganic and organic salts. Accordingly, regardless of the disclosure in Merck, there is no specific disclosure of the use of sodium or potassium chloride in the claimed amount in Lion ‘775. For this reason, the present claims are certainly not anticipated by Lion ‘775, either taken by itself or read in view of Merck.

Moreover, even the combined teachings of Lion ‘775 and Merck cannot make the present claims obvious. Specifically, there is nothing in either of these references which would even remotely suggest the advantages provided by the presently claimed compositions.

In support of the assertion that the presently claimed compositions provide unexpected benefits, Applicants direct the Examiner’s attention to the results already of record in the present specification. In particular, Table 1 provides stability results for a number of compositions, which either contain no additive or one of various additives. For the Examiner’s convenience, Table 1 is reproduced below.

¹ See, the entries for “electrolyte,” in Hackh’s Chemical Dictionary, Fourth Edition, McGraw-Hill, New York, pp. 234-235, 1969; and Concise Chemical and Technical Dictionary, Fourth Edition, Chemical Publishing, New York, p. 478, 1986, copies of which are attached hereto, as Exhibits A and B, respectively.

Table 1

Examples	Additives	25 °C		
		immediately after dissolution	one day later	7 days later
Comparative Example 1	no additive	×	×	×
Comparative Example 2	sodium sulfate	×	×	×
Comparative Example 3	sodium dihydrogen phosphate	×	×	×
Example 1	sodium chloride	○	○	○
Example 2	potassium chloride	○	○	○
Example 3	sodium lactate	○	○	○
Example 4	sodium glutamate monohydrate	○	○	○
Example 5	POE (40) hardened castor oil pyroglutamate isostearate diester	○	○	○

Assessment: dissolved transparently: ○; precipitated: ×.

As can be seen from the results presented in Table 1, the compositions which contained either sodium chloride or potassium chloride exhibited improved stability as compared to analogous compositions which contained a different inorganic electrolyte such as sodium sulfate or sodium dihydrogen phosphate.

In addition, Tables 2 and 3 report the results of stability testing for a number of compositions, which contain vary amounts of either sodium chloride or potassium chloride. Again, for the Examiner's convenience, Tables 2 and 3 are reproduced below.

Table 2

Examples	Amount of sodium chloride added	25 °C		
		immediately after dissolution	one day later	7 days later
Comparative Example 1	not added	×	×	×
Example 6	0.05%	○	Δ	Δ
Example 7	0.1%	○	○	○
Example 8	0.2%	○	○	○
Example 9	0.5%	○	○	○
Example 10	1.0%	○	○	○

Assessment: dissolved transparently: ○; slightly precipitated: Δ; precipitated: ×.

Table 3

Examples	Amount of potassium chloride added	25 °C		
		immediately after dissolution	one day later	7 days later
Comparative Example 1	not added	×	×	×
Example 11	0.05%	Δ	Δ	Δ
Example 12	0.1%	○	○	○
Example 13	0.2%	○	○	○
Example 14	0.5%	○	○	○
Example 15	1.0%	○	○	○

Assessment: dissolved transparently: ○; slightly precipitated: Δ; precipitated: ×.

As can be seen from the results presented in Tables 2 and 3, the addition of either sodium or potassium chloride is effective for stabilizing the composition. Applicants submit that there is nothing in either of the cited references which would remotely suggest the improved stability afforded by the use of sodium or potassium chloride as compared to the use of other inorganic electrolytes such as sodium sulfate or sodium dihydrogen phosphate. Accordingly, these results are truly unexpected and ensure the patentability of the present claims.

For these reasons, the cited references cannot make the present claims obvious, and the rejection should be withdrawn.

Applicants submit that the present application is now ready for examination on the merits, and early notification of such action is earnestly solicited.

Respectfully submitted,

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